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7.B Mole and Mass Avogadro's Number, The Mole, Grams, Atoms, Molar Mass Calculations - Introduction Converting Between Grams and Moles Mole Ratio Practice Problems Concept of Mole - Part 1 | Atoms and Molecules | Don't Memorise Empirical Formula Molecular Formula Determination From Percent Composition Molality Practice Problems - Molarity, Mass Percent, and Density of Solution Examples ~~Introduction to Moles~~ Stoichiometry Basic Introduction, Mole to Mole, Grams to Grams, Mole Ratio

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Practice Problems Mole Conversions Made Easy: How to Convert Between Grams and Moles How To Convert Grams To Moles - VERY EASY! Counting Atoms: Intro to Moles Part 2 ~~Converting Grams to Moles Using Molar Mass | How to Pass Chemistry The Mole: Avogadro's Number and Stoichiometry Introduction to Combustion Analysis, Empirical Formula /u0026 Molecular Formula Problems Dalton's Law of Partial Pressure Problems /u0026 Examples - Chemistry Stoichiometry Mole to Mole Conversions—Molar Ratio Practice Problems Step by Step Stoichiometry Practice Problems | How to Pass Chemistry Balancing Chemical Equations Practice Problems Concept of Mole | Avogadro's Number | Atoms and Molecules | Don't Memorise Chemistry B Moles Packet Answer~~

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PAGE 9 CHEMISTRY WORKSHEET # 6 MIXED MOLE PROBLEMS (GRAMS, MOLECULES, AND LITERS) You now know three things a mole can be: a molar mass, 6.02×10^{23} molecules and, for a gas, 22.4 liters at STP.

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PAGE 4 CHEMISTRY WORKSHEET # 2: THE MOLE AS A UNIT OF MASS Now that you know how to find the mass of one mole of a substance (molar mass) you can easily find the mass of several moles or the mass of a fraction of a mole using the factor-label technique.

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A mole of carbon contains 6.02×10^{23} atoms of carbon. A mole of helium contains 6.02×10^{23} atoms of helium. A mole of sodium contains 6.02×10^{23} atoms of sodium. A mole of gold contains 6.02×10^{23} atoms of gold. However, we have learned about seven elements that exist as diatomic molecules— H_2 , N_2 , O_2 , F_2 , Cl_2 , Br_2 , and I_2

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PAGE 4CHEMISTRY WORKSHEET # 2: THE MOLE AS A UNIT OF MASS
Now that you know how to find the mass of one mole of a substance (molar mass) you can easily find the mass of several moles or the mass of a fraction of a mole using the factor-label technique. 1 mol = a molar mass of an atom/molecule (g/mol)
EXAMPLE: What is the mass of 5.00 moles of water H₂O?
STEP 1: 2 H = 2 x (1.01) = 2.02
STEP 2: ?

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grams H₂O = 5.00 moles H₂O x 18.02 g H₂O = 90.10 ...

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PAGE 6 CHEMISTRY WORKSHEET # 4 MOLAR MASSES AND

AVOGADRO ' S NUMBER Now that you know two definitions

of a mole (a molar mass and an Avogadro ' s number of particles) you can combine these two definitions into one

problem. EXAMPLE: How many molecules are there in 90.1

grams of water? 2 H = 2 x (1.01) = 2.02 O = 1 x (16.00) =

16.00 H₂O = 18.02 g? molecules H₂O = 90.1 g H₂O x 1

mole H₂O x 6.02 x 10²³ molecules H₂O = 3.01 x 10²⁴

molecules H₂ ...

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PAGE 9 CHEMISTRY WORKSHEET # 6 MIXED MOLE

PROBLEMS (GRAMS, MOLECULES, AND LITERS) You now

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CHEMISTRY WORKSHEET #5 MOLE PROBLEMS-MOLAR
VOLUME OF A GAS We Have Learned That A Mole Is A Mass
Of Material And Number Of Particles. A Mole Can Also Be A

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Measure Of Volume When We Are Talking About Gases. You May Remember From Previous Science Classes That All Gases Behave Basically The Same As Far As The Physical ...

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